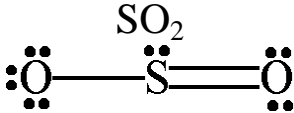


## ANSWERS: Polarity of molecules

$\text{CBr}_4$ (or $\text{CCl}_4$ ) non-polar	$\text{CH}_3\text{Br}$ polar
$\text{CH}_2\text{Cl}_2$ polar	$\text{CHCl}_3$ polar
$\text{H}_2\text{S}$ polar	$\text{CO}_2$ non-polar
$\text{CH}_3\text{Cl}$ polar	$\text{NH}_3$ polar
$\text{CF}_4$ non-polar	$\text{OCl}_2$ polar
$\text{COCl}_2$ polar	$\text{SiH}_4$ non-polar
$\text{SO}_2$  polar <p>please note the more accepted Lewis diagram for <math>\text{SO}_2</math> is an expanded octet, this is Level 3 Chemistry</p>	$\text{NCl}_3$ polar